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6 Chemical Reactions That Changed History

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Answer Key 1. What is organic chemistry? The branch of chemistry dealing with compounds of carbon. 2. How many carbon compounds exist? Millions 3. How many bonds can carbon form? Four bonds 4. When was nylon first developed? Why was its development so essential? Before World War II. A substitute for silk. 5. What was the first organic synthesis?

The World of Chemistry

World of Chemistry Copyright Houghton Mifflin Company. All rights reserved. b. Cl₂(g) + 2KI(aq) 2KCl(aq) + I₂(s) 0.125 mol KI 2 mol KCl 2 mol KI = 0.125 mol KCl 0.125 mol KI 1 mol I₂ 2 mol KI = 0.0625 mol I₂ c. Na₂B₄O₇(s) + H₂SO₄(aq) + 5H₂O(l) 4H₃BO₃(s) + Na₂SO₄(aq) 0.125 mol Na₂B₄O₇ 4 mol H₃BO₃ 1 mol Na₂B₄O₇ = 0.500 mol H₃BO₃ 0.125 mol Na₂B₄O₇

Chapter 9 Chemical Quantities

World of Chemistry Copyright Houghton Mifflin Company. All rights reserved. 20. a. mass of 1 mol Ba = 137.3 g = 137.3 g mass of 2 mol Cl = 2 (35.45 g) = 70.90 g mass of 8 mol O = 8 (16.00 g) = 128.0 g molar mass of Ba(ClO₄)₂ = 336.2 g b. mass of 1 mol Mg = 24.31 g = 24.31 g mass of 1 mol S = 32.07 g = 32.07 g

Chapter 6 Chemical Composition

Answer Key 1. Why is it important to use the correct amount of materials in a chemical reaction? If too little is used the reaction may not proceed very far. The use of too much chemical may result in waste. 2. What names are given to the materials at the beginning and end of a chemical reaction? Reactants and products. 3.

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together by chemical bonds. Carbon dioxide and water are two examples, but answers will vary. 3. Rutherford's experiment involved shooting a beam of particles at a thin sheet of metal foil. According to the then-current "plum pudding" model of the atom, most of these positively charged particles should have passed right through the foil.

Chapter 3: Standard Review Worksheet

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