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Koretsky Department of Chemical Engineering Oregon State University

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or  $\ln \frac{P_2}{P_1} = \frac{h_{vap}}{R} \left( \frac{1}{T_1} - \frac{1}{T_2} \right)$

where  $P_1 = 101 \text{ kPa}$ ,  $T_1 = 373 \text{ K}$ ,  $T_2 = 373 \text{ K}$

so  $\ln \frac{P_2}{101} = \frac{h_{vap}}{R} \left( \frac{1}{373} - \frac{1}{T_2} \right)$

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(a) The Clausius-Clapeyron equation:  $dP_i^{\text{sat}}/P_i^{\text{sat}} = \Delta h_i^{\text{vap}}/RT^2$  or  $\ln P_i^{\text{sat}}/101 \text{ kPa} [ ] \# \$ \% \& ' = ( ) \Delta h_i^{\text{vap}}/R (1/373 \text{ [K]}^* + , - . /$  so  $P_i^{\text{sat}} = 101 \text{ kPa} [ ] ( ) \exp \# \Delta h_i^{\text{vap}}/R (1/373 \text{ [K]} \$ \% \& ' ( )^* + , - .$

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