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which the state variables of a thermodynamic system are changed.

(Examples: pressure changed, or temperature changed, or both changed but with some other condition ...)

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Questions. 13. 1 kg of liquid iron will have greater entropy, since it is less ordered than solid iron and its molecules have more thermal motion. In addition, heat must be added to solid iron to melt it; the addition of heat will increase the entropy

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of the iron. 15.

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The Conservation of Energy law allows energy to flow bi-directionally between its various forms. For example in a pendulum, energy continually goes to/from kinetic energy and potential energy. Entropy is different: No conservation law – the entropy

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19 Then Joshua said to Achan, “ My son, I implore you, give glory to the Lord, the God of Israel, and give praise to Him; and tell me now what you have done. Do not hide it from me. ” 20 So Achan answered

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Joshua and said, “ Truly, I have sinned against the Lord, the God of Israel, and this is what I did. ” .

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ENTROPY AND THE SECOND LAW
OF THERMODYNAMICS. 1. In a
reversible process the system: A. is always

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Close to equilibrium states B. is close to equilibrium states only at the beginning and end C. might never be close to any equilibrium state D. is close to equilibrium states throughout, except at the beginning and end E. is none of the above ans: A 2.

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continually goes to/from kinetic energy
and potential energy.

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Thermodynamics-4/19/11 The basic idea
of the 2nd Law of Thermodynamics is that

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not all energy processes are equally likely. Even though they all conserve energy, nature “ prefers ” certain types of processes over others.

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However, for a system moving to

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Thermodynamics, $\Delta S_{univ} > 0$, because the Second Law states that for any spontaneous process, the entropy of the universe increases. 20-4 20.32 Since entropy is a state function, the entropy changes can be found by summing the entropies of the products and subtracting the sum of the entropies of the reactants.

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Questions. 1. Yes, mechanical energy can be transformed completely into heat or internal energy, as when an object moving over a surface is brought to rest by friction. All of the original mechanical energy is converted into heat.

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